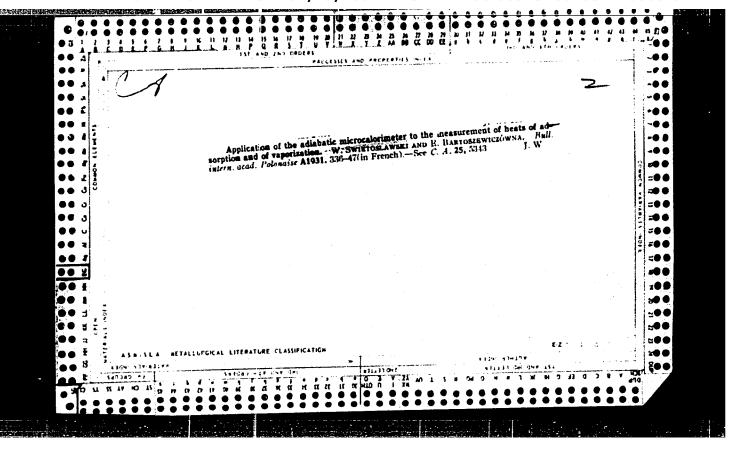
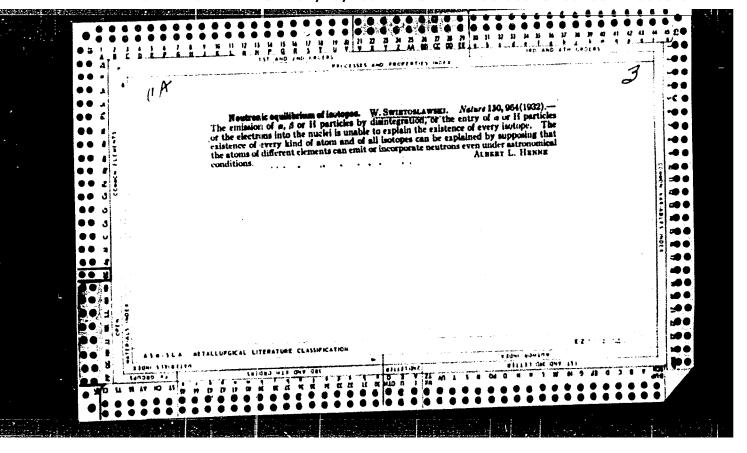
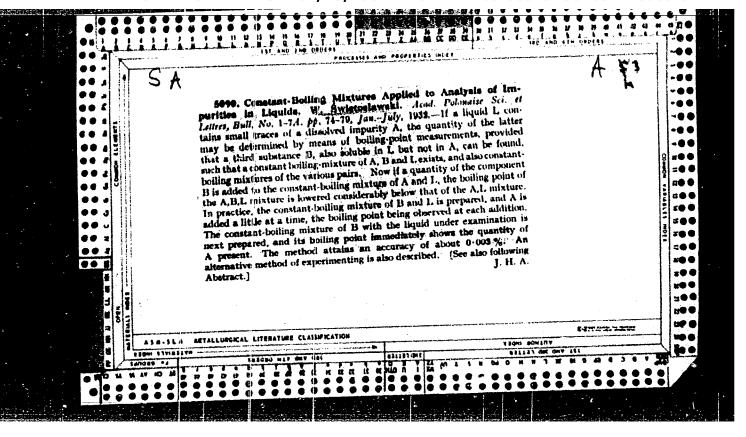
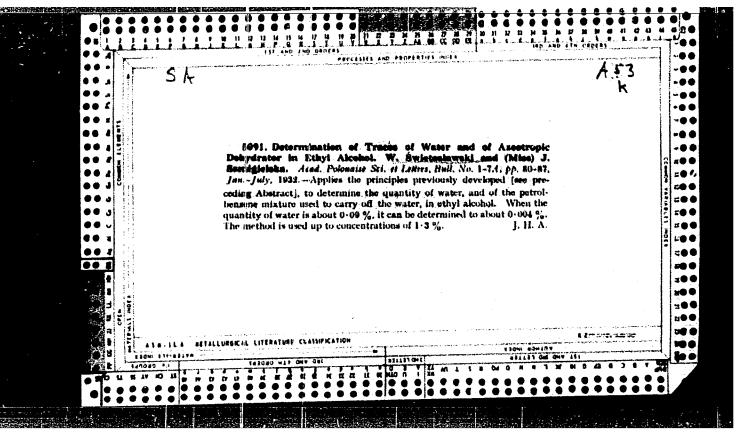


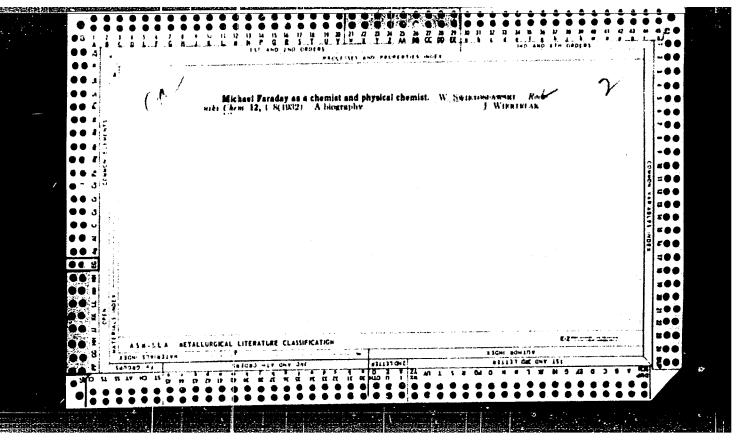
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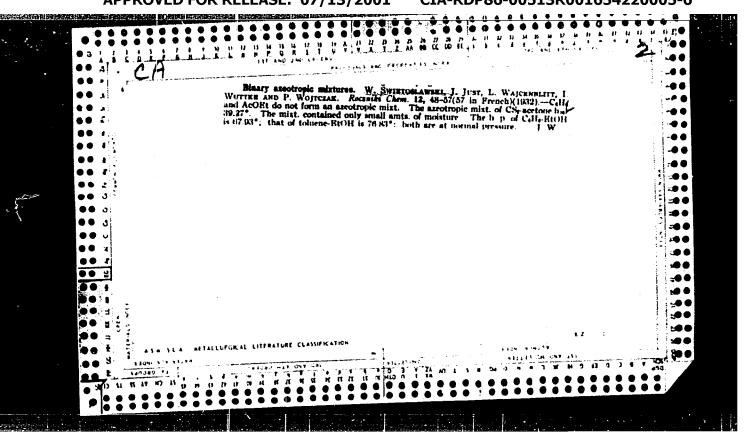


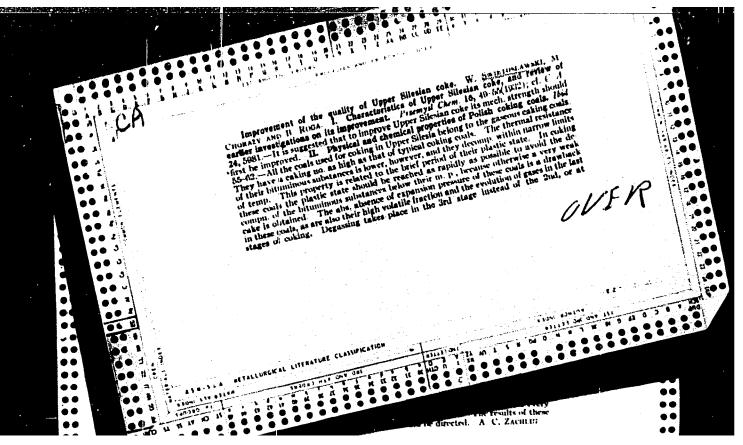


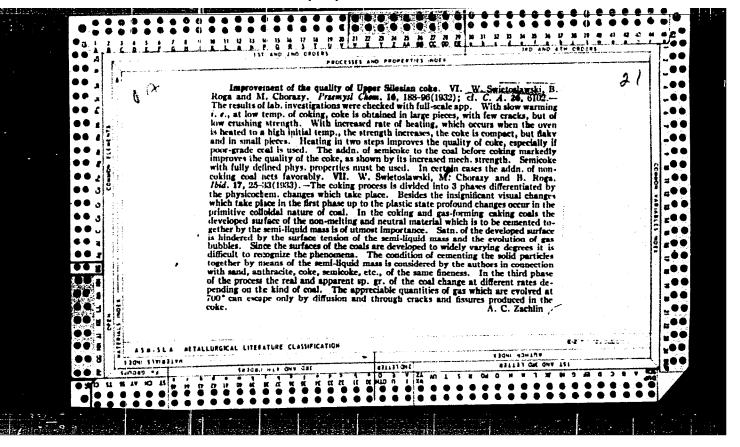


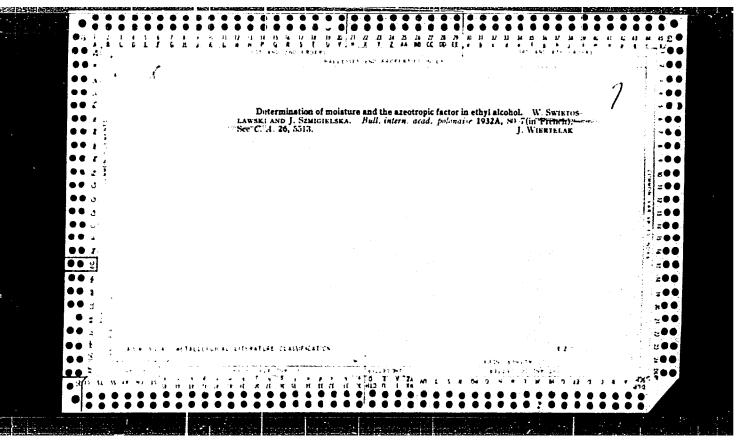


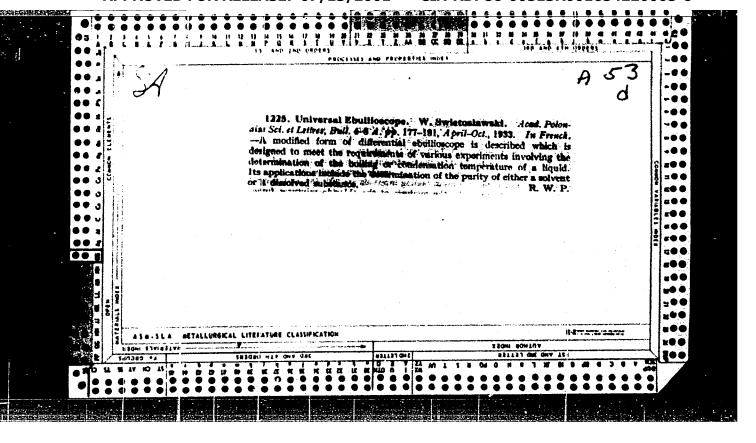


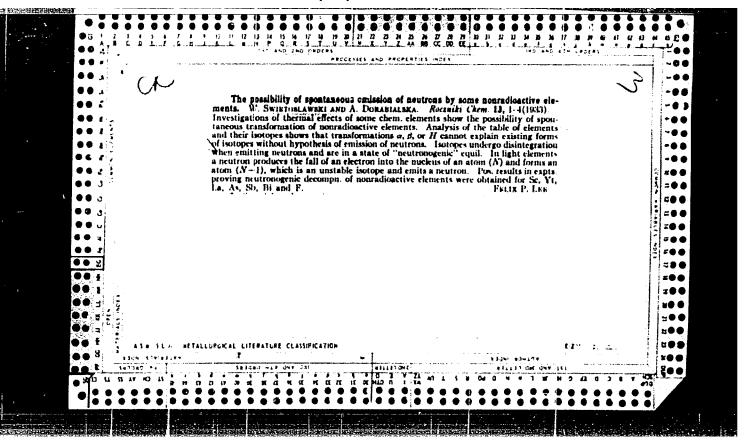


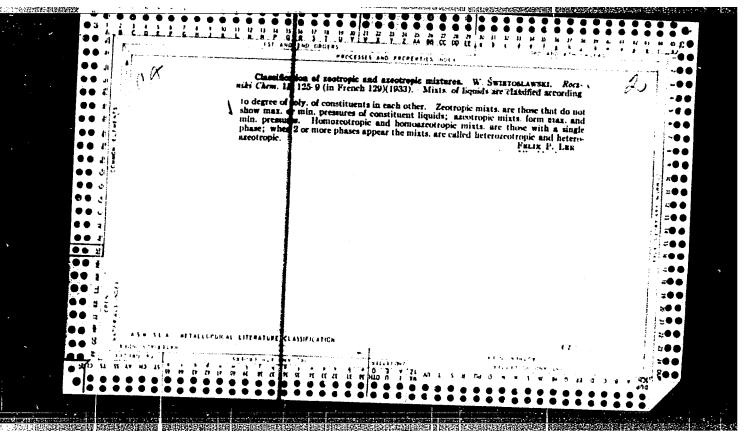


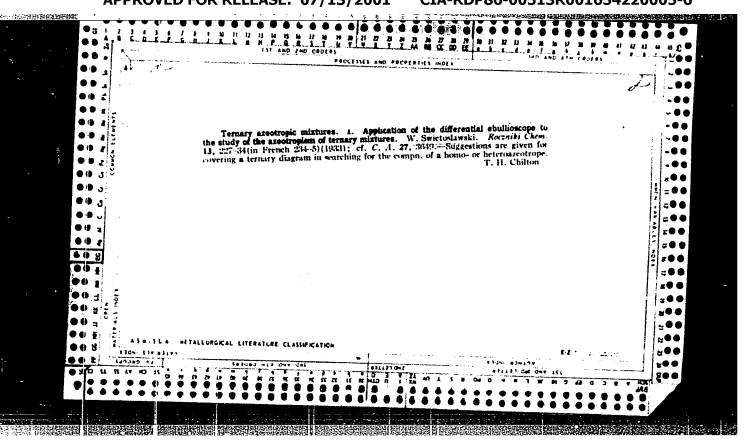


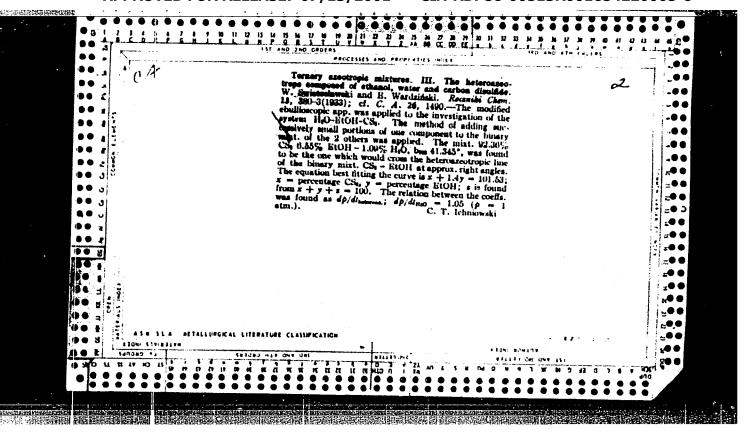


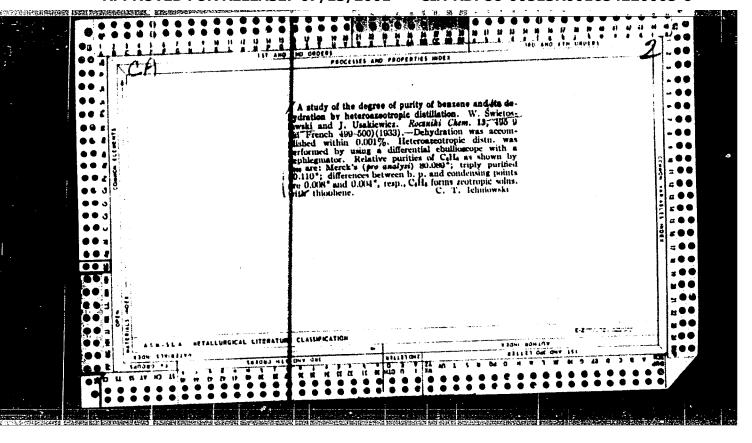


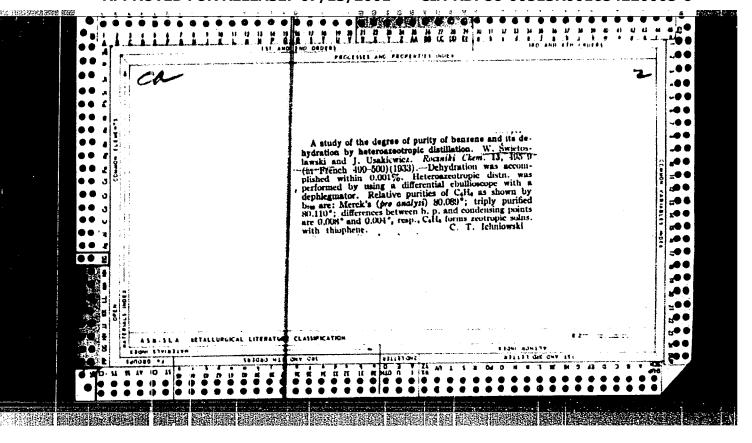


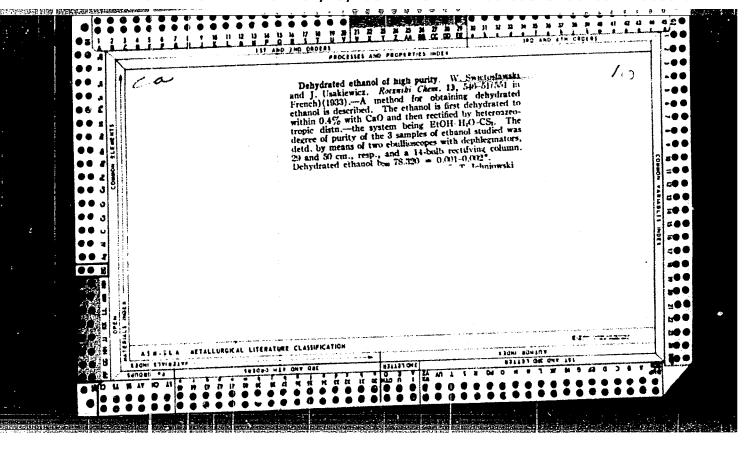


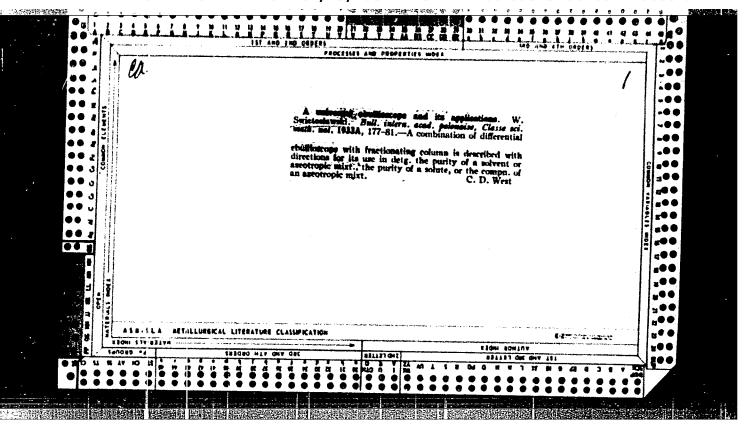


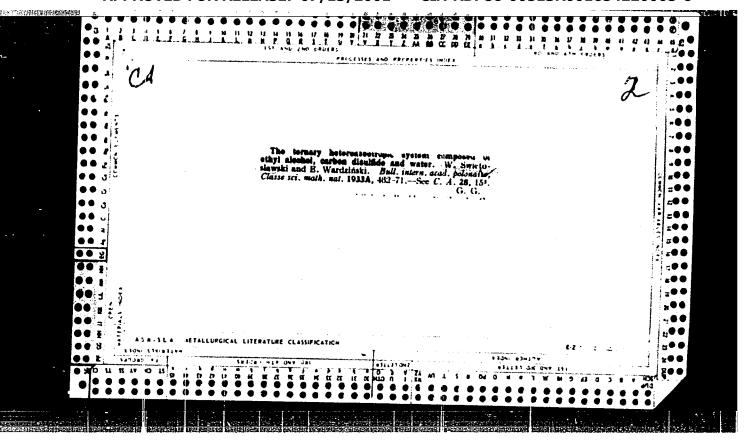


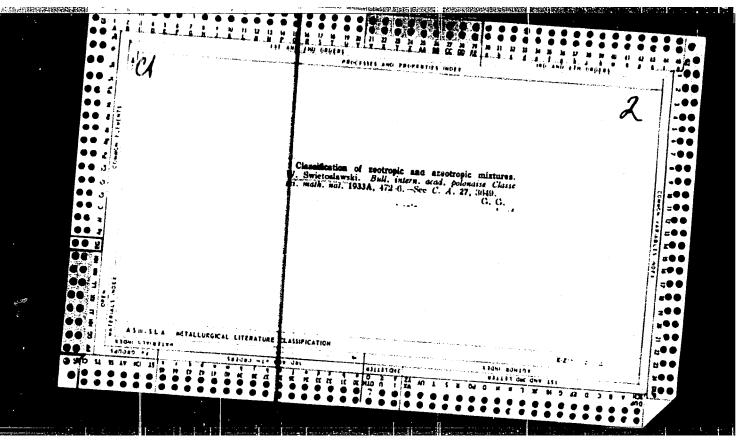


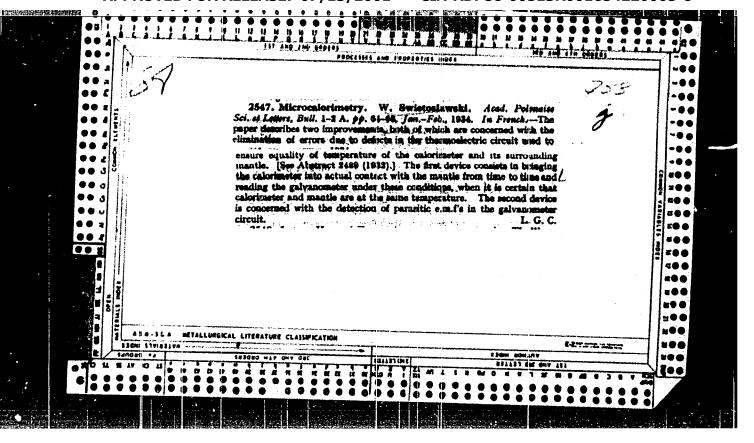




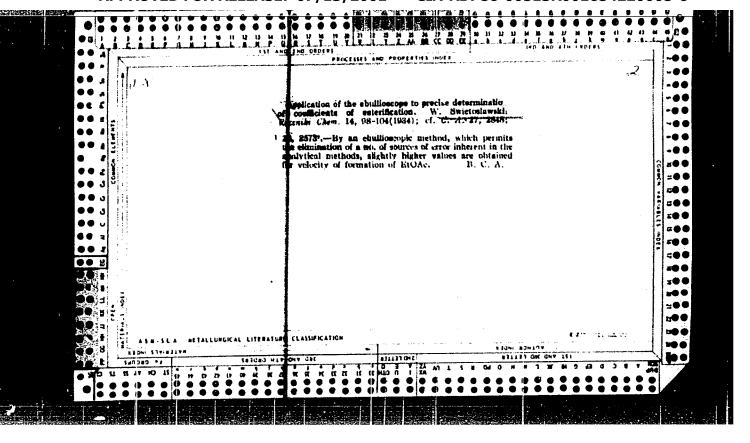


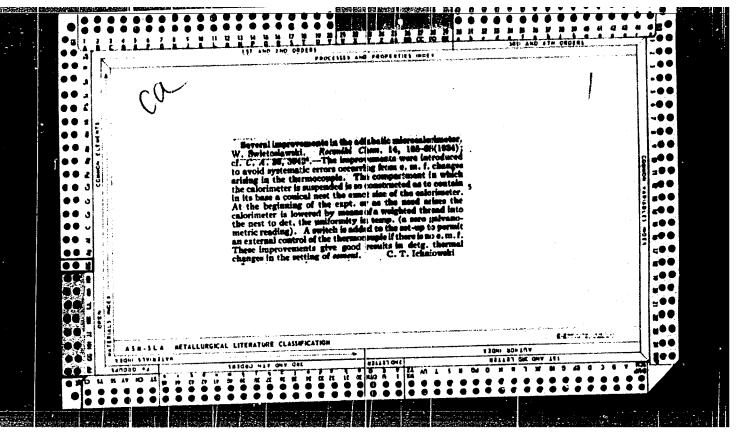


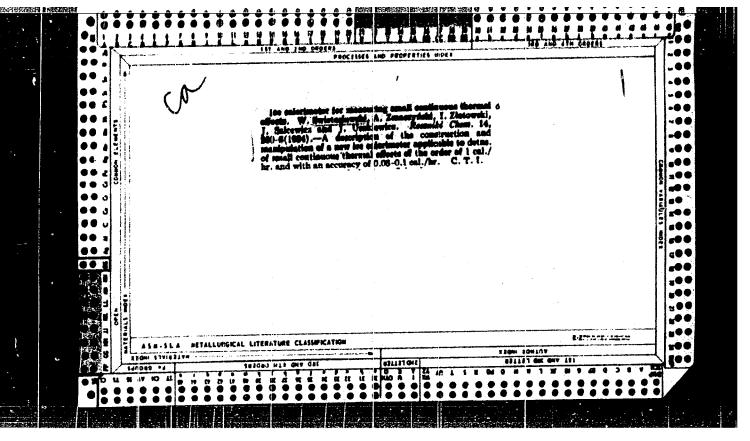


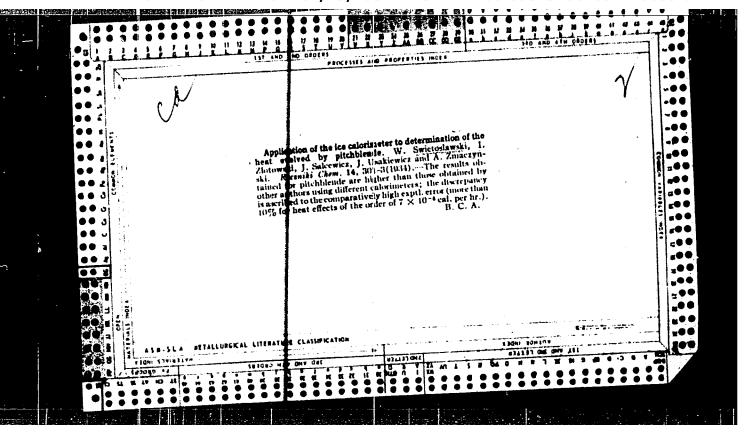


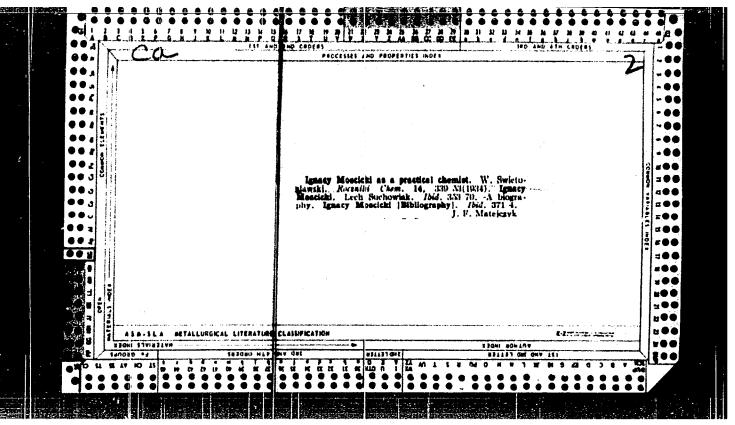
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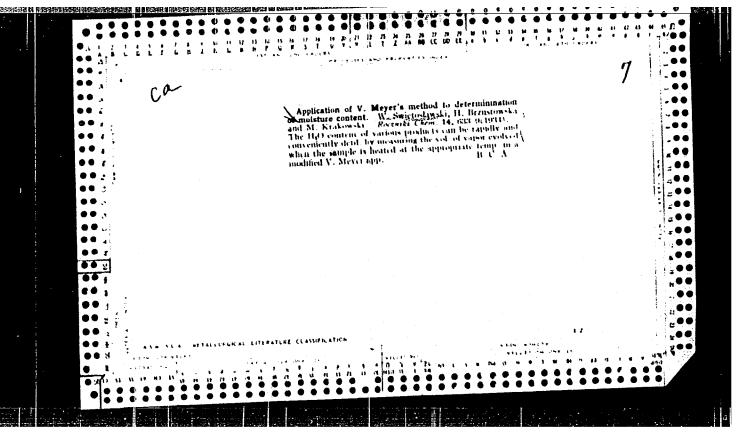


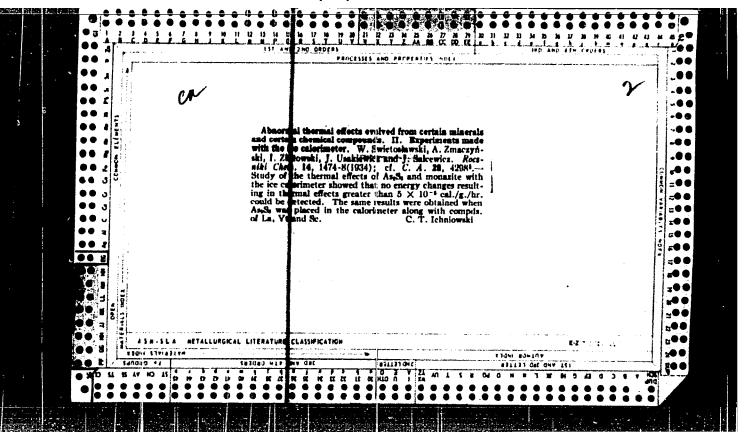


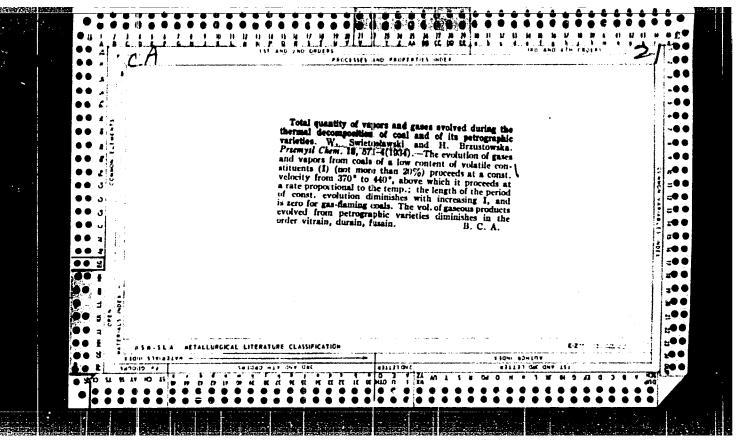


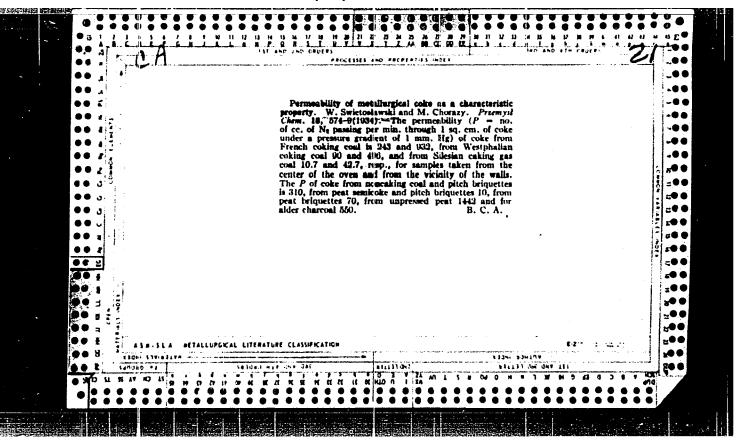


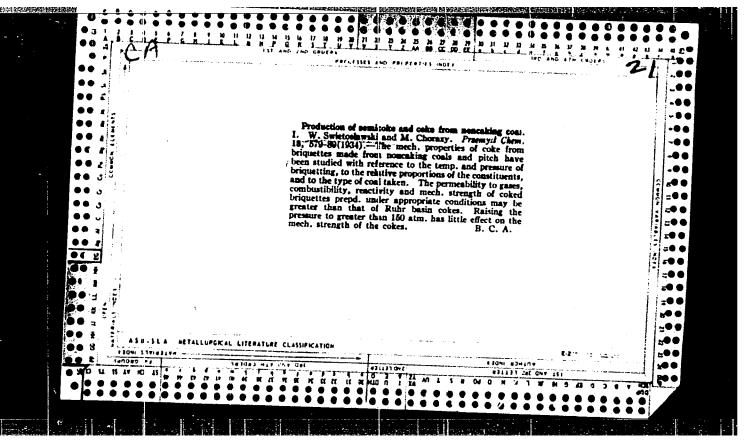


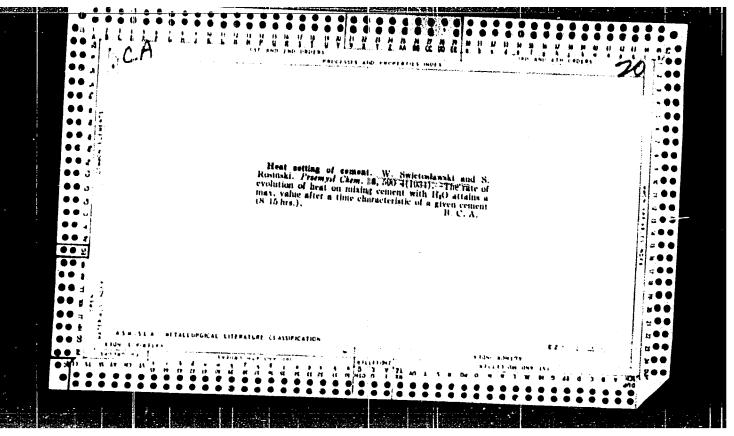


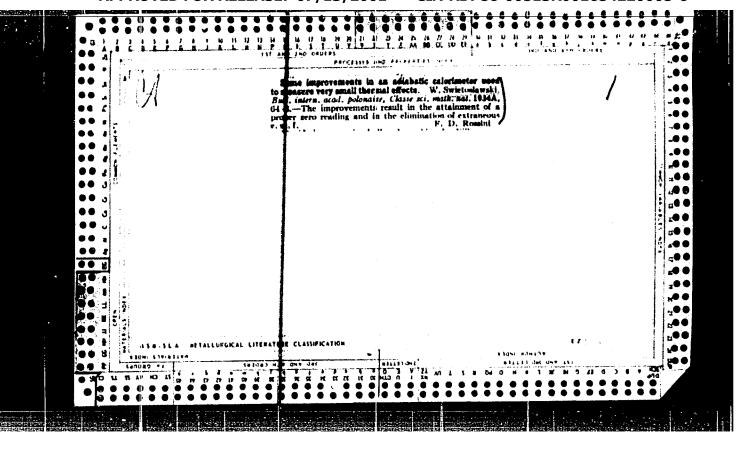


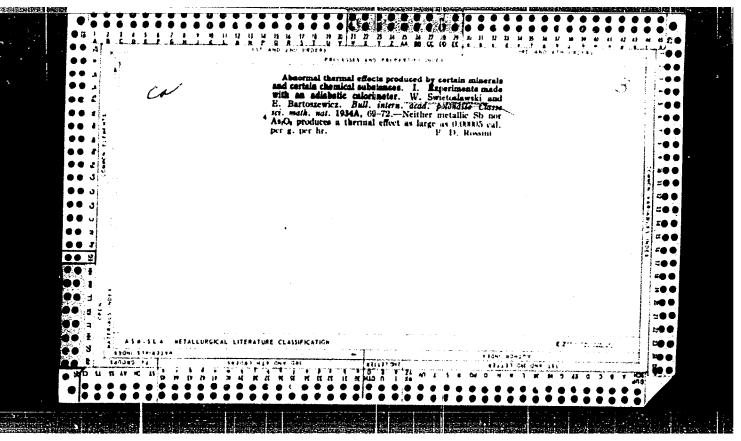


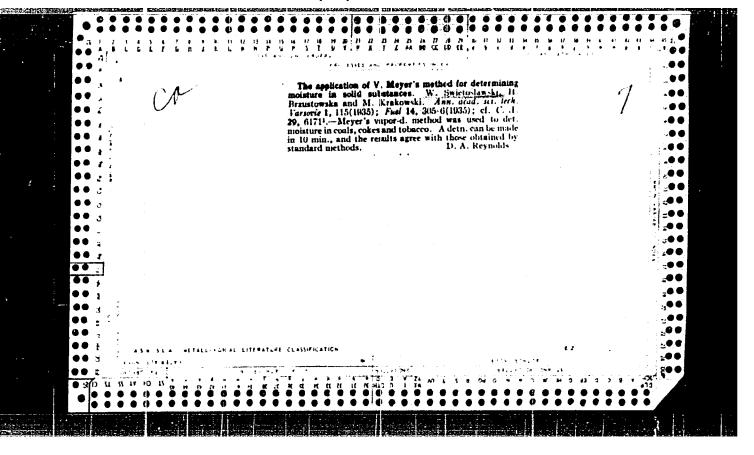


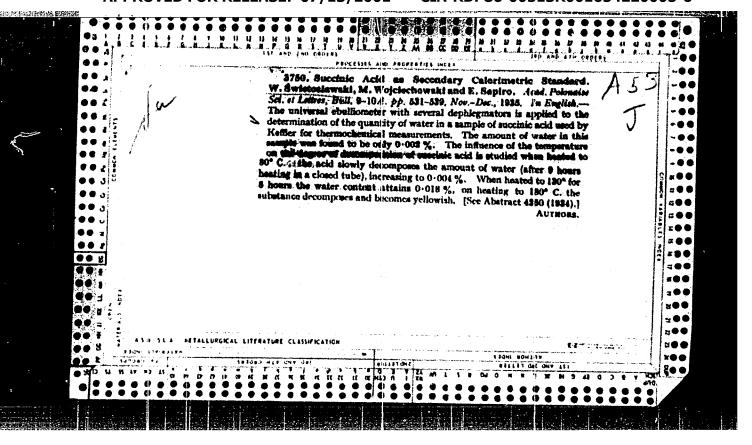


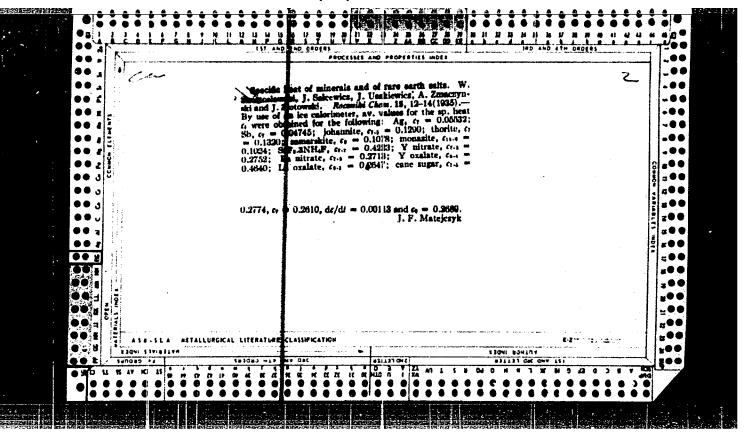


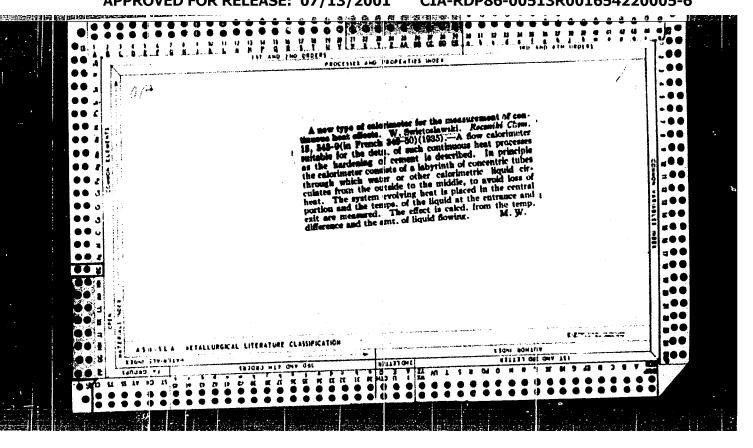


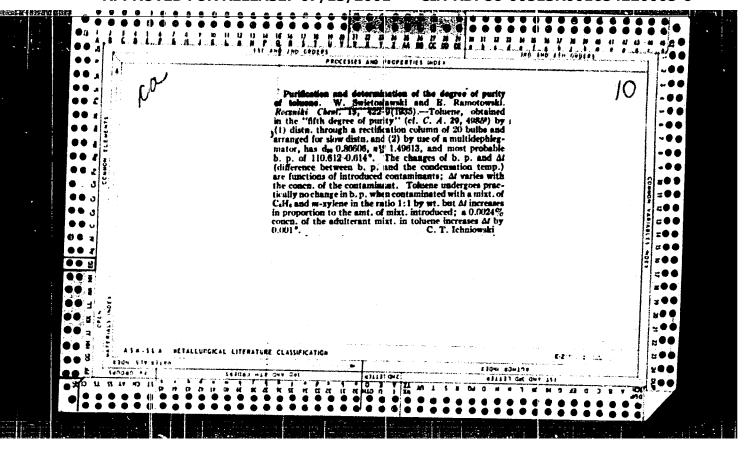


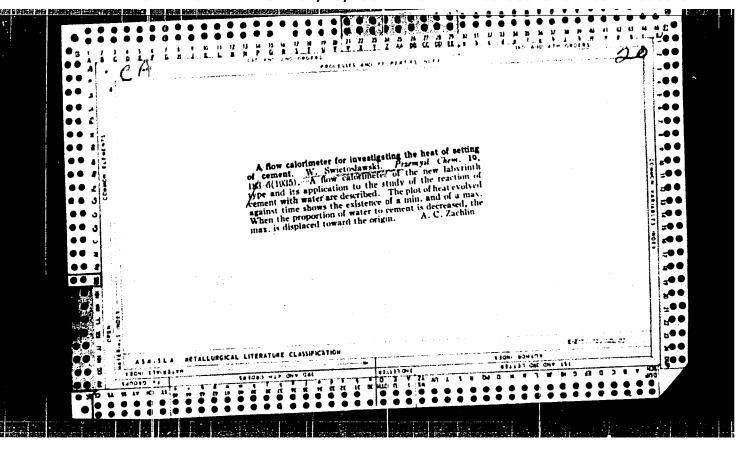


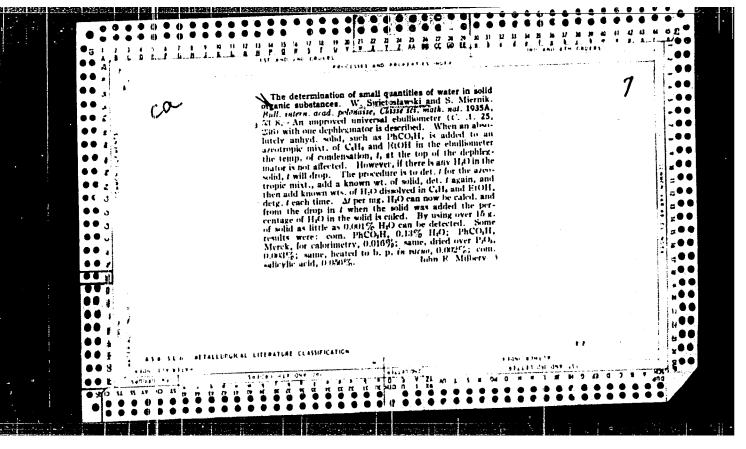


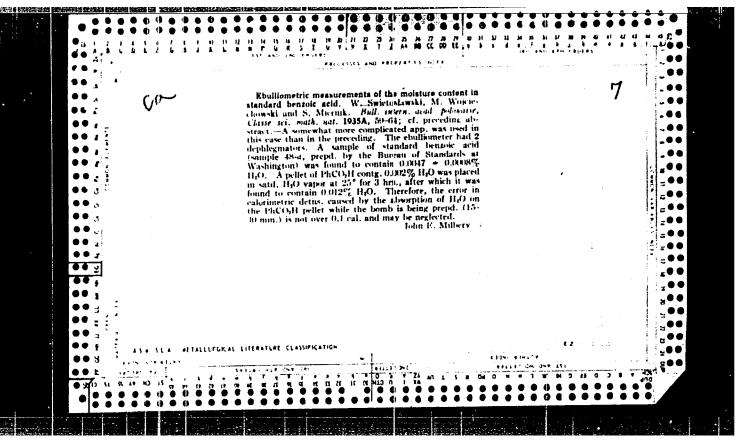


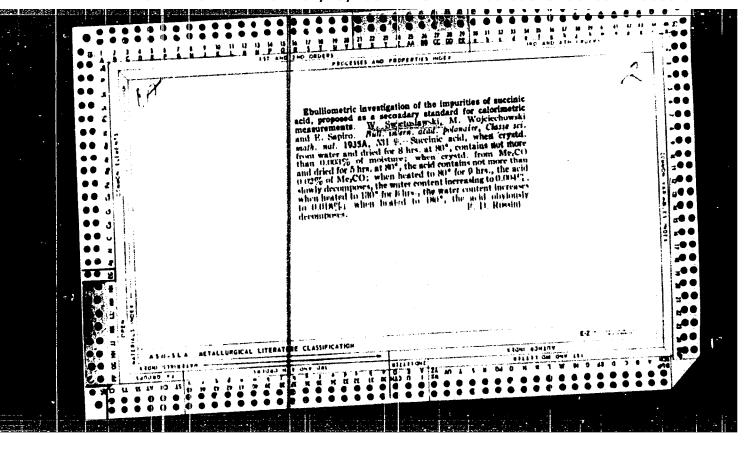


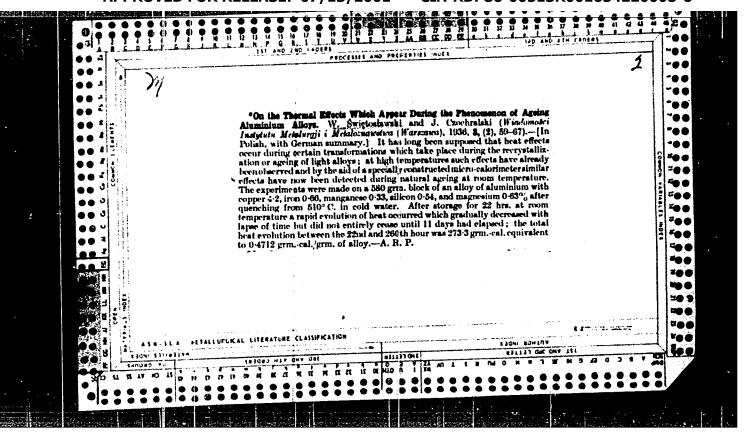


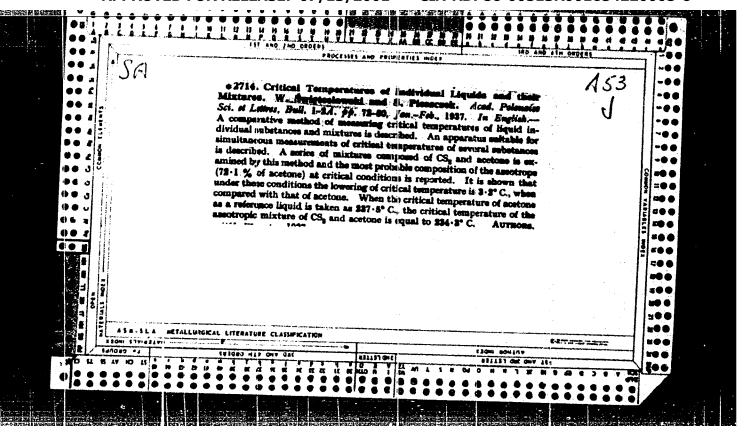


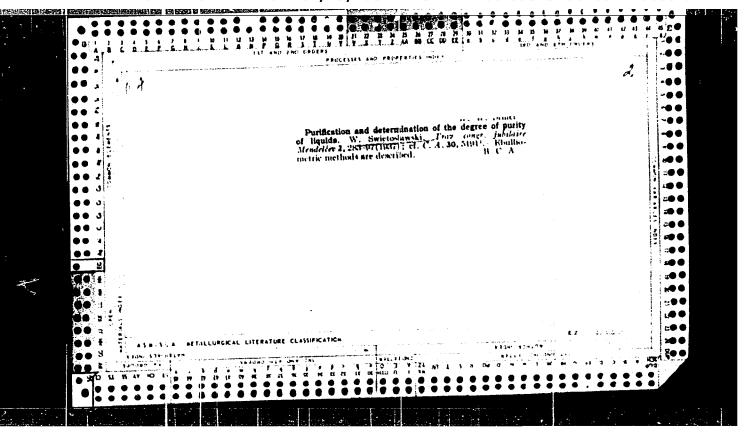


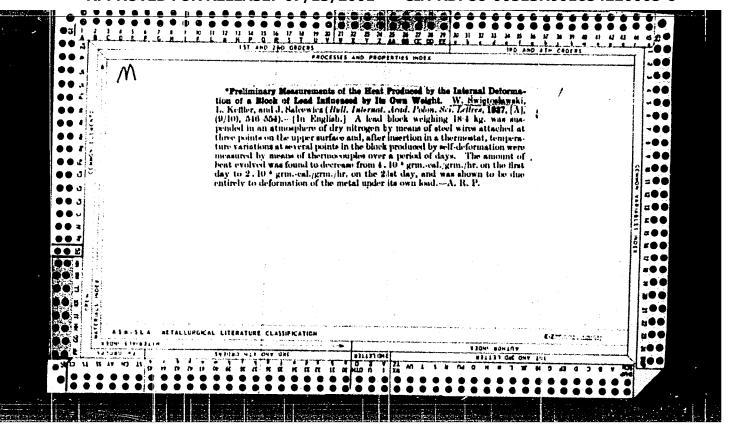


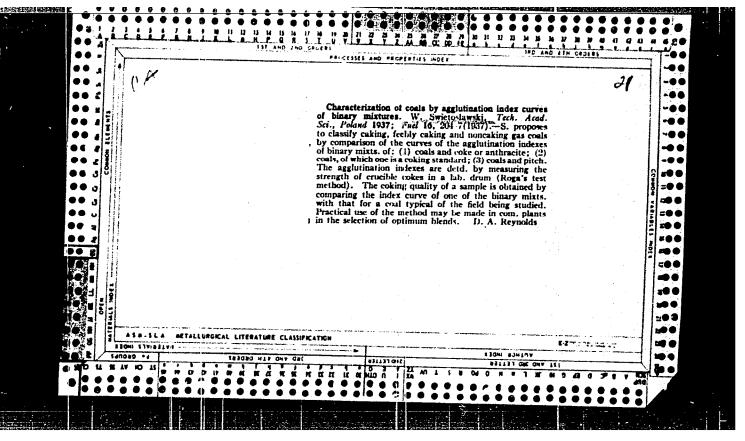


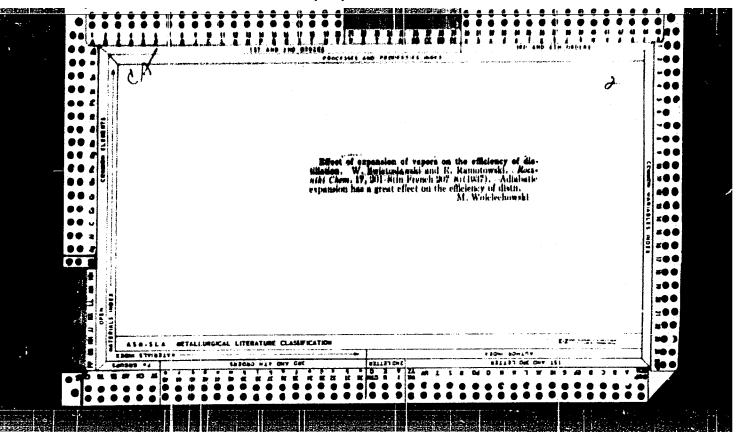


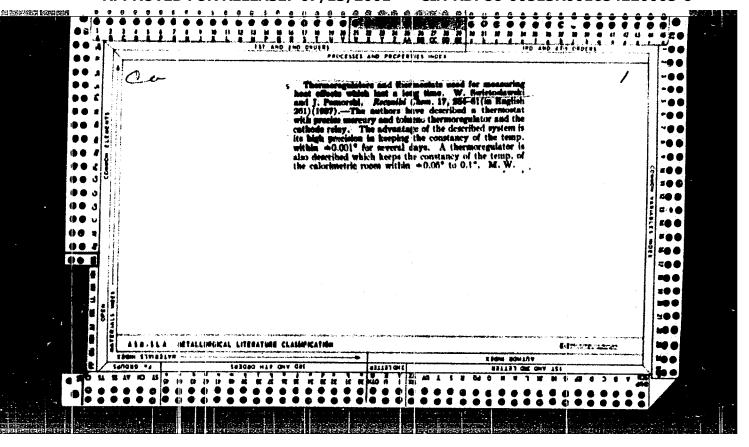


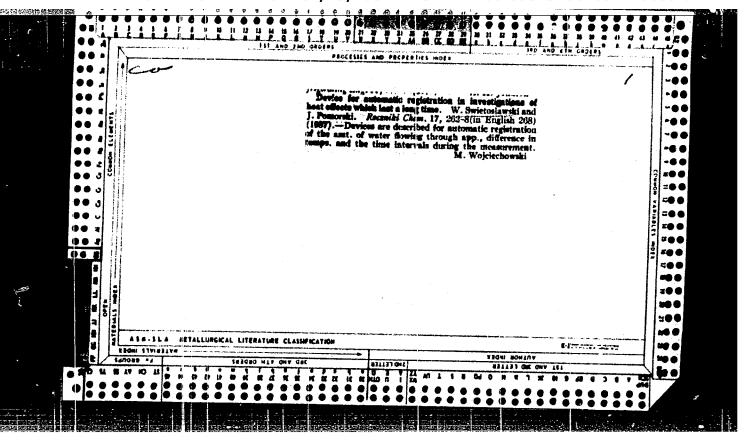


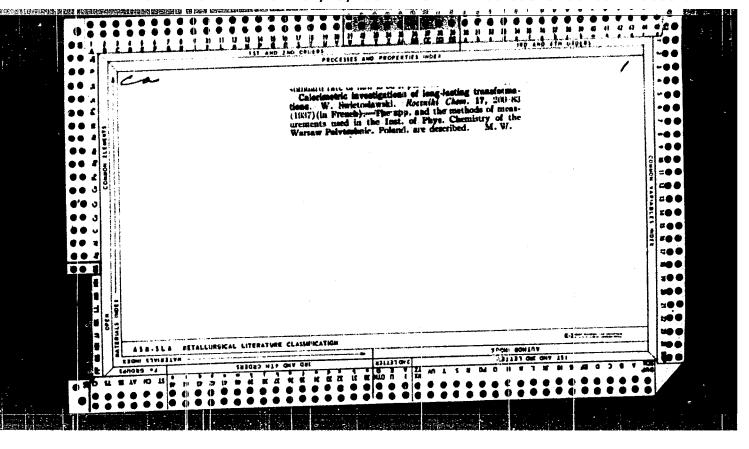


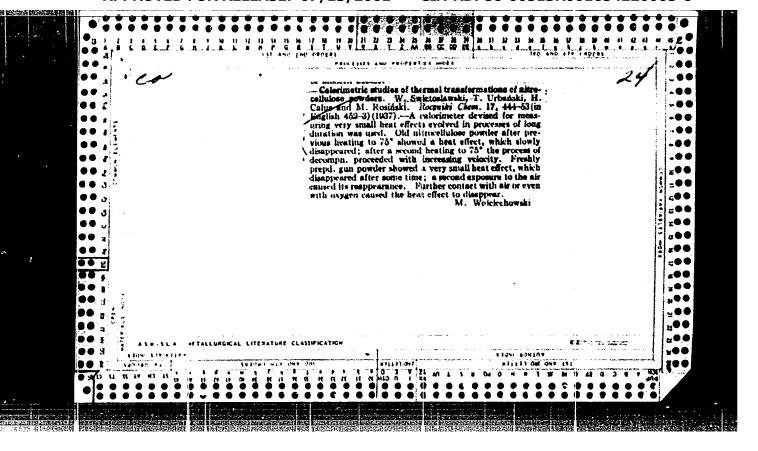


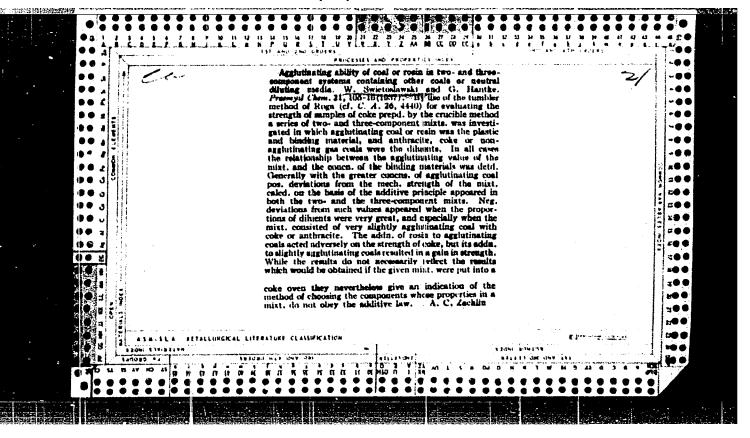


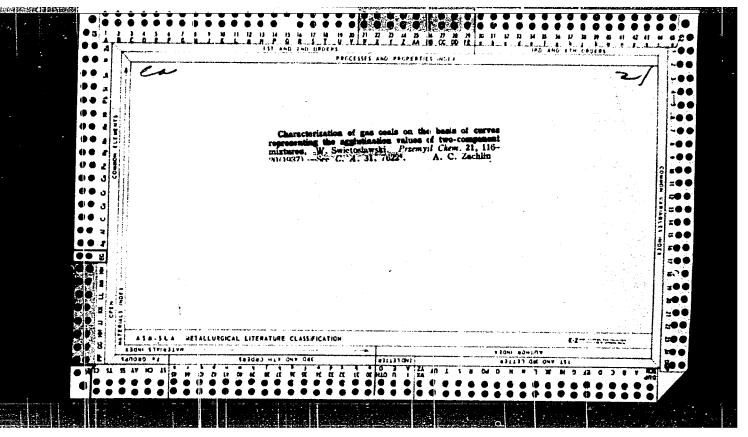


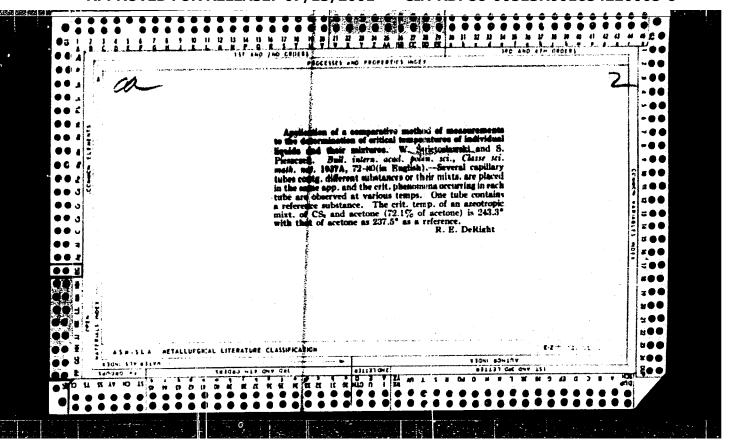


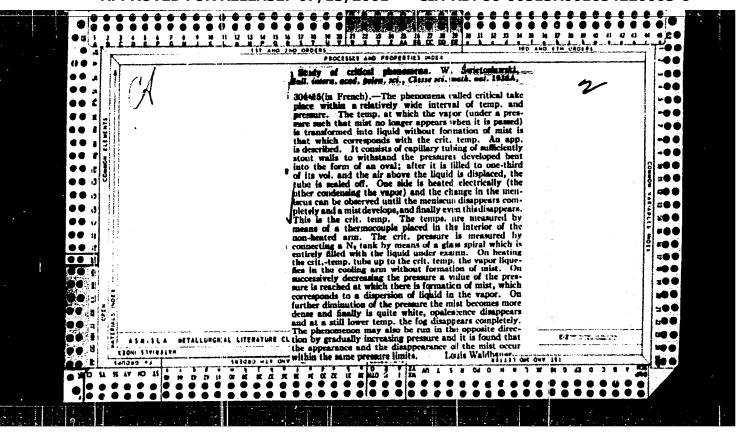


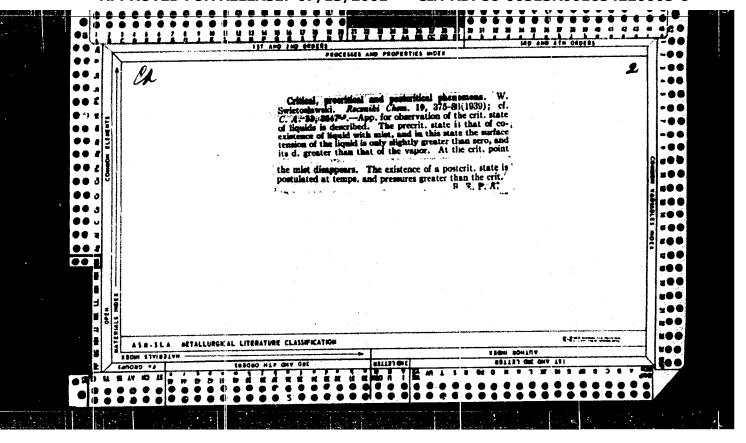


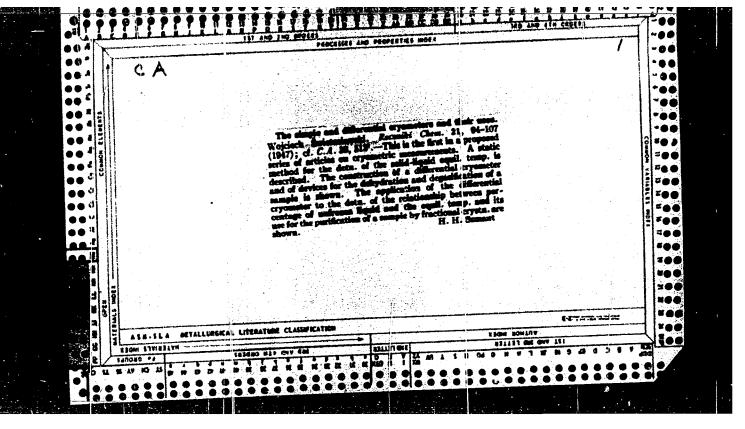


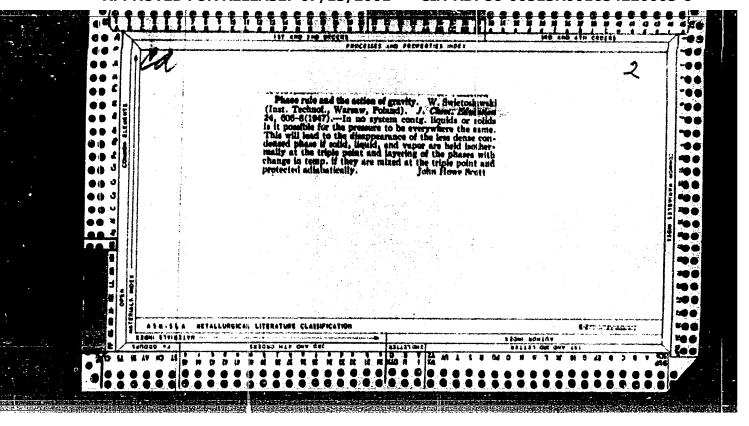


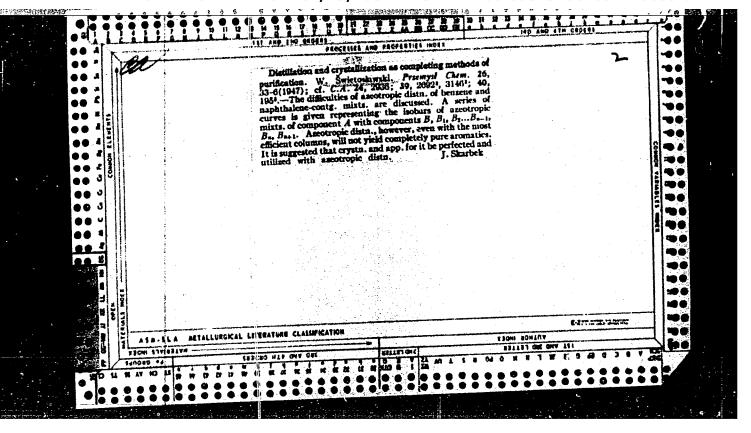


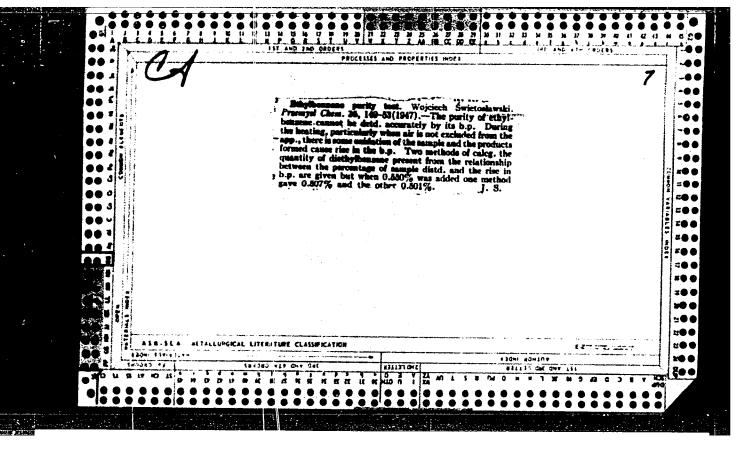


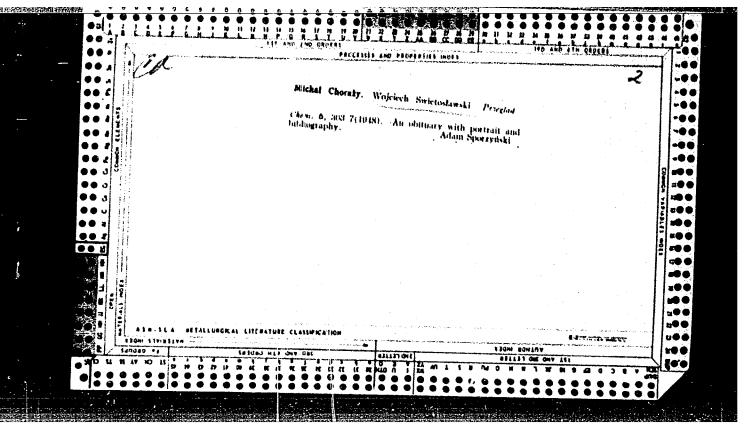


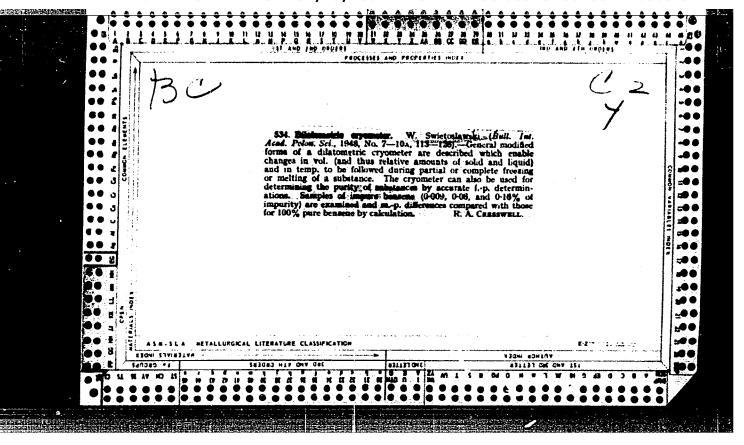


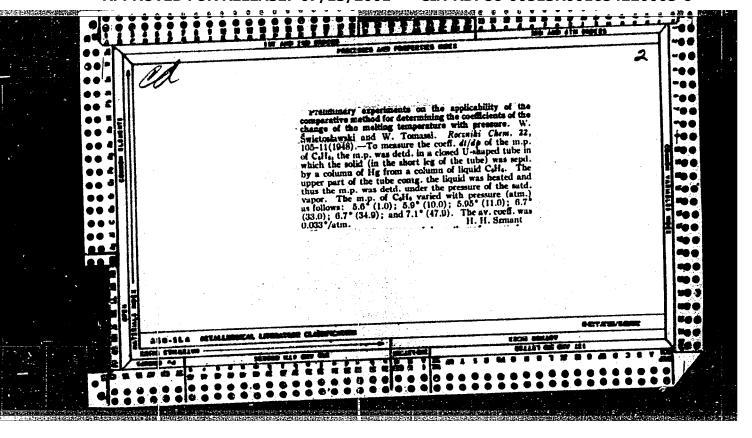




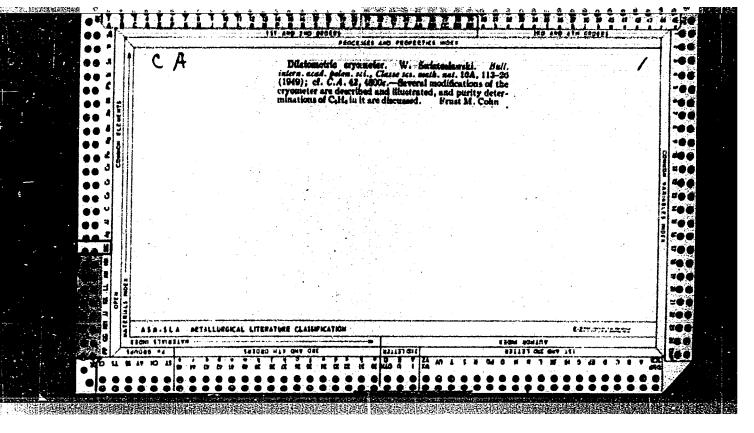


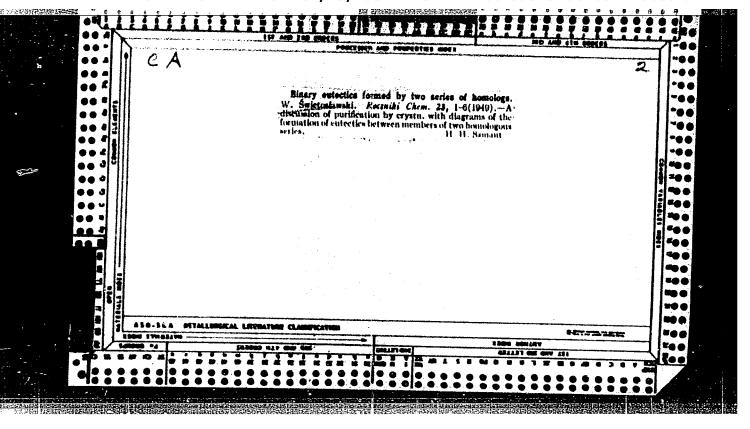


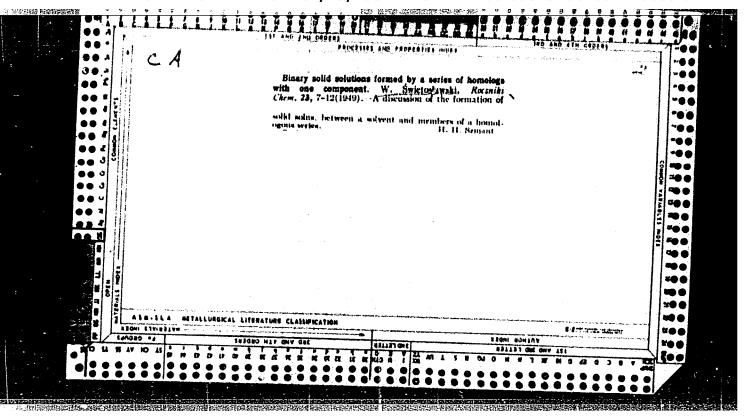




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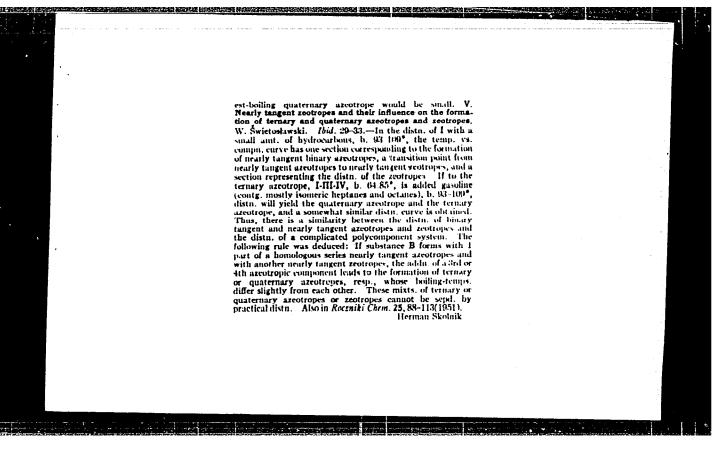
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The quaternary azeotrope n-heptane-benzene-ethaned-water. I. W. Świetosławski, and K. Zieborak (Central Inst. Ind. Chem. Research, Warsaw). Bull. intern. acad. polon. sci., Classe sci. math., et nat. Ser. A. 1990, 9-12(in English).—The quaternary azeotrope, C.Ha[I]-C;Ha[II]-EtOH(III)-Hr(O(IV), b. 64.79°, contained 62.4 I, 12.1 II, 18.7 III, and 6.8 wt.—70° IV and consisted of 2 phases: the lower, die 0.8772, no 1.3789, was 17.2 vol.—70 or 17.85 wt.—70° at 20° of the total and contained 11.8 I, 0.9 II, 54.4 III, and 32.9 wt.—70° IV; the upper, die 0.8385, no 1.54.4 III, and 32.9 wt.—70° IV; the upper, die 0.8385, no 1.54.4 III, and 32.9 wt.—70° IV; the upper, die 0.8385, no 1.54.4 III, and 32.9 wt.—70° IV; the upper, die 0.8385, no 1.54.4 III, and ding small amts. of the higher-boiling a differential challiometer with the lower-boiling ternary azeotrope (I-III-IV) and adding small amts. of the higher-boiling ternary azeotrope (I-III-IV) and against the compan. of the mixt. The quaternary azeotrope benzene-ethanol-water-isofctane. Ibid. 13-14.—The quaternary azeotrope benzene-ethanol-water-isofctane. Ibid. 13-14.—The quaternary azeotrope benzene-ethanol-water-isofctane. (V)-Ibid. 13-14.—The quaternary azeotrope. I-isofctane (V)-Ibid. 13-14.—The quaternary azeotrope in India II.5 I. 1.2 V.54.6 III, and 6.7 wt.—70° IV and consisted of 2 phases: the lower, die 0.8766, wt.—70° IV and consisted of 2 phases: the lower, die 0.8766, wt.—70° IV and consisted of 2 phases: the lower, die 0.8766, wt.—70° IV, III, and 6.7 liv. the upper, die 0.8203, ng 1.4595, contained 72.2 I, 17.0 V., 9.9 III, and 0.9 wt.—70° IV. III. The quaternary azeotrope composed of benzene, ethanol, water, and cyclowase. K. Zieborak. Ibid. 15-18.—The quaternary azeotrope, I-III-IV-cyclohexane (VI), b. 62.10°, contained 54.0 VI, 21.5 I, 17.4 III, and 7.1 wt.—70° IV. IV. Tangent and formal proposed and proposed appropriate assetropes. W. Swietcoshawkii. Ibid. 19-28.—By use of the isobar curves of an azeotropic lower. Acuternary agent isobars limiting th

the azeotropic range is defined as the extreme b.p. limits of the corresponding homologs which form tangent or nearly tangent isobars (i.e., the upper and lower limits of azeotropy). The formation of ternary azeotropes of A and C with a series of homologs, B. B., B., etc., or their isomers, depends upon the smaller azeotropic range of A with B's, although the ternary azeotrope range may be somewhat larger than the smaller binary azeotrope range may be somewhat larger than the smaller binary azeotrope range may be somewhat larger than the smaller binary azeotrope range. The formation of quaternary azeotropes is limited by the azeotropic capacity of the binary systems (A with B'g, h., ving the smallest range; all 3 agents (A, C, and D) should form azeotropes with each other and with the series of homologs (B, B<sub>1</sub>, B<sub>2</sub>, etc.) within a certain range; the quaternary azeotrope range may be somewhat larger than the smallest binary range by virtue of nearly tangent isobars of the binary system. In a similar manner, it is concluded that a 5-component azeotrope might exist, although the probability of such formation is small, and its isolation would be difficult since the azeotropic depression with respect to the low-



Swith Tosz Hosky do

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Swietoslawski W. "On the Classification of Sorbents and Ion-Exchangers." (W sprawie klasyfikacji sorbentow i jonitow). <u>Przemysl Chemiczny</u>, No. 1, 1950, pp. 41-43.

The author suggests a classification of sorbents and ion-exchangers into: perfect and imperfect groups. To the first belong all sorbents and ionites with one single function, while the other comports those which, in addition to being sorbents also play the part of ionites and vice versa. A suggestion is made to divide ionites into one-, two-, three-, and multi-function groups. The swelling of ionites interferes to some degree with the proposed grouping. There follows a characterization of the secondary processes occuring on non-ideal sorbents and ion-exchangers. A number of Polish terms are proposed for practical use.

SO: Polish Technical Abstracts - No. 2, 1951

